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Explanation, Numericals *Guided Notes The Atom*

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iTeachly.com 3 above comes after the H. This means there are two H's (hydrogen atoms) in each molecule of H<sub>2</sub>O. 2. Coefficients – the regular-sized numbers that tell you how many molecules you have. For example, in 3H<sub>2</sub>O, the 3 is the coefficient. This shows that there are three H<sub>2</sub>O molecules in

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Guided Notes: The Atom History of the Atom

Democritus- Greek philosopher; 400 BC 1st to use word “atomos” (means \_\_\_\_\_) Aristotle-refuted what Democritus said; all matter comes from “hyle” Early Alchemists-1200-1500’s; tried to turn material into \_\_\_\_\_; kept chemistry alive;

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Guided Notes: Atomic Structure 01 Atom: the smallest particle of an element that retains its \_\_\_\_\_ in a \_\_\_\_\_. Democritus believed that atoms are \_\_\_\_\_ & \_\_\_\_\_

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Think about it: Was Democritus correct? Two issues with Democritus's ideas: 1. His ideas did not explain \_\_\_\_\_.

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Guided Notes on Thomson's Model of the Atom  
Original Theories of Matter: 1) Democritus: matter is made up of indivisible particles called atoms 2) Dalton: elements, compounds, law of conservation of mass Which of the above theories of matter needs to be revised to become more accurate now that we know the results of Thomson's experiments with electrons?



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The Structure of the Atom Glossary TERM DEFINITION  
anion an ion that has a negative charge atom atomic  
mass atomic weight the average atomic mass of a  
naturally occurring sample of an element cation  
compound a substance made up of two or more  
chemical elements that are chemically bonded  
together electron element energy level ion isotope  
mass number the total number of protons and  
neutrons in an atom neutron a subatomic particle that  
has about the same mass as a proton but does not  
have an ...

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the atom remained electrically neutral (i.e. had the same number of negative and positive charges); a

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neutral particle would also have to be present. This led to the discovery and measurement of the mass of the neutron by Chadwick in 1932, completing the picture of the subatomic particles which makes up the atom.

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## *Atomic Structure Guided Notes Answers*

1. \_\_\_ The atomic number is equal to this particle of the atom. 2. \_\_\_ This particle has no charge. 3. \_\_\_ These make up the mass of an atom. 4. \_\_\_ This particle has a negative charge. 5. \_\_\_ These are found in the nucleus of an atom. Use the student responses as a formative assessment to determine if further instruction on atoms is warranted.

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Unit 4B- Electron Configuration- Guided Notes Atomic Structure Electrons are arranged in \_\_\_ or \_\_\_ around the nucleus of an atom o First shell can hold a

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maximum of \_\_\_\_\_ electrons o Second shell can hold a maximum of \_\_\_\_\_ electrons o Third shell can hold a maximum of \_\_\_\_\_ electrons ...

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Reveals the links between an atom's structure and its chemical destiny showing how an atom makes its passage through nature.

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